

## Chapter 8 Covalent Bonding Test A Answers

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CHAPTER 8 SOLUTIONS MANUAL Covalent Bonding Covalent Bonding Solutions Manual Chemistry: Matter and Change • Chapter 8 121 Section 8.1 The Covalent Bond pages 240–247 Practice Problems page 244 Draw the Lewis structure for each molecule. 1.  $\text{PH}_3$   $\text{H}-\text{H}-\text{H}$   $\text{H}-\text{H}-\text{H}$   $\text{H}-\text{H}-\text{H}$  respectively, for single, double, and triple P — — 2.  $\text{H}_2\text{S}$   $\text{H}-\text{H}-\text{H}$   $\text{H}-\text{H}-\text{H}$  —  $\text{H}-\text{S}-\text{S}$  ...

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Chapter 8 Covalent Bonding and Molecular Structure 8-6 Example:  $\text{PF}_3$  Example:  $\text{CO}_2$  Step 1: Count Valence Electrons Count the total number of valence electrons in the molecule or ion. Anions have extra electrons, so add 1 electron for each negative charge. Cations have a deficiency of electrons, so subtract 1 electron for each positive charge.

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8.8: Strengths of Covalent Bonds Bond enthalpy is the enthalpy change for the breaking of a particular bond in one mole of a gaseous substance o The designation  $D(\text{bond type})$  is use to represent bond enthalpies Many important bonds such as C—H bond only exist in polyatomic molecule

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In covalent bonding, atoms still want to achieve a noble gas configuration (the octet rule). 12. In covalent bonding, atoms still want to achieve a noble gas configuration (the octet rule). But rather than losing or gaining electrons, atoms now share an electron pair. 13. In covalent bonding, atoms still want to achieve a noble gas ...

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