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The pressure of a gas is directly proportional to the temperature if

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the volume remains constant What is the formula for the combined gas law? $P_1V_1/T_1 = P_2V_2/T_2$

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answer choices to indicate what each unit of measurement is used to measure. A. Pressure B. Volume C 1. K 4. kPa A 2. atm 5. L 3. mL 6. °C C. Temperature 'A 7. A 8.

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volume, and pressure of a gas when the number of particles is constant freezing point of water in Fahrenheit and Celcius 32 degrees F, 0 degrees C

chapter 3 section 3.2 THE GAS LAWS Flashcards | Quizlet

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The van der Waals equation is $(P + \frac{a}{V^2})(V - nb) = nRT$ where

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P = pressure V = volume a = pressure correction constant unique to the gas b = volume correction constant unique to the gas n = the number of moles of gas T = absolute temperature The van der Waals equation includes a pressure and volume correction to take into account the interactions between molecules. Unlike ideal gases, the individual particles of a real gas have interactions with each other and have definite volume.

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The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon, as it forms around particles of a metal catalyst. Carbon nanotubes are a recently discovered allotrope of carbon. Three other allotropes of carbon—buckyballs, graphite, and diamond—are illustrated at the left, as is the molecule methane, CH_4 , from which nanotubes and buckyballs can

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be made. The element carbon forms an amazing number of compounds with structures that follow from simple methane, found in natural gas, to the complex macromolecules that serve as the basis of life on our planet. The study of chemistry also follows from the simple to the more complex, and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement.

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